

# TutorBot: Your 24/7 Chemistry Tutor

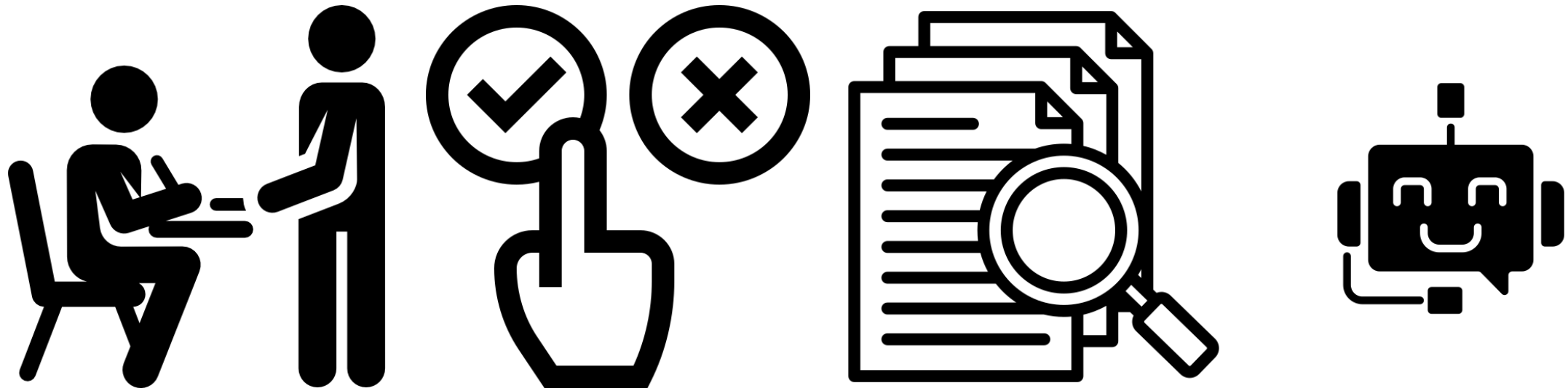
Standalone Web App for Student  
Support

Tom Fuller, Chris Collison

## Tutor Bot



- Build courses with assignments and questions
- Create AI assistants specific to your course
- Capture student interactions with the AI
- Get feedback from the students on the AI's performance



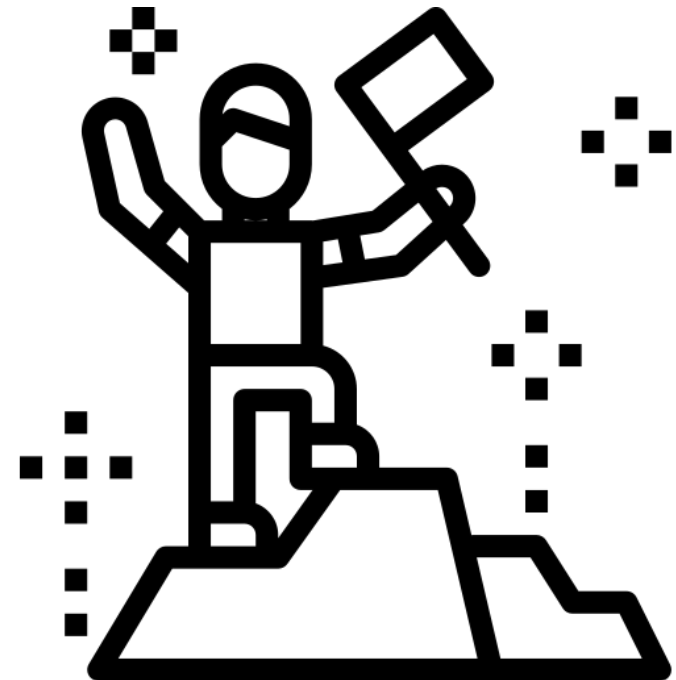
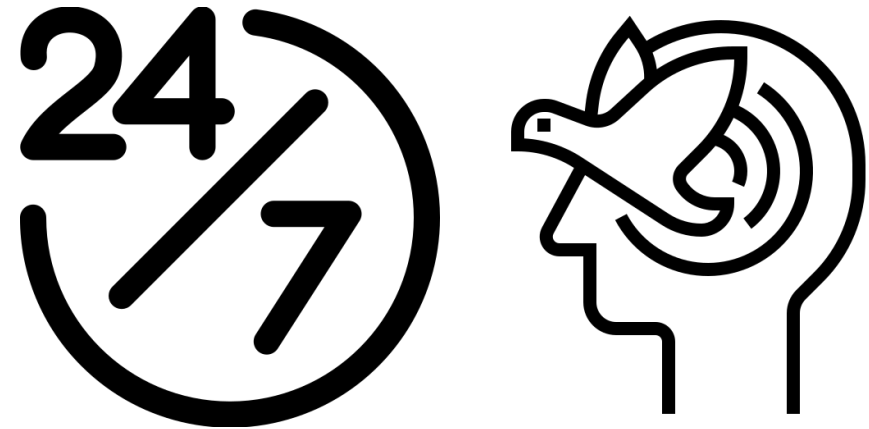
## TutorBot Workflow

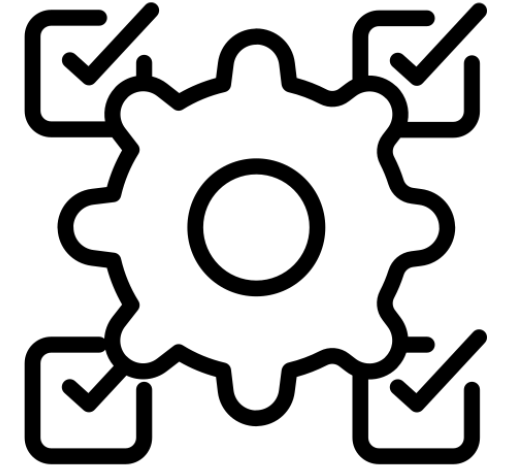
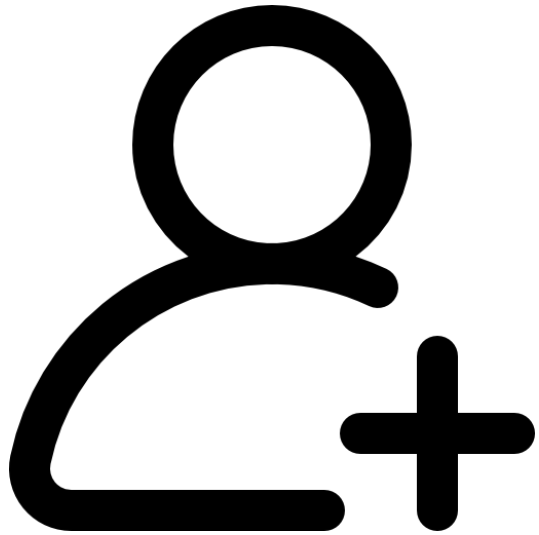
- Write a student question and define concepts
- Provide a correct answer example and common mistakes
- Upload reference materials (textbook PDFs, notes)
- Define the bot's role and tasks

# Why TutorBot?

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- Supports independent problem-solving
- Accessible 24/7
- Encourages critical thinking and concept mastery





## Breakout Session Goals

- Sign up for OpenAI API account
- Gain access to TutorBot as Faculty
- Learn to write an active question for your class
- Receive a checklist for implementation of a pilot

# Join us



TutorBot

My Chats

My Courses

My Assistants

Christopher Collison



CHEM-171



BobBob



Exam 1



limiting\_reactant



## Question

A chemist combines 28.74 grams of nitrogen gas (N<sub>2</sub>) with 12.46 grams of hydrogen gas (H<sub>2</sub>) in an attempt to produce ammonia (NH<sub>3</sub>). Which reactant is the limiting reagent, and how much ammonia can theoretically be produced from this mixture? Molar masses for Nitrogen gas, Hydrogen gas and ammonia gas are respectively: 28.02 g/mol, 2.016 g/mol and 17.03 g/mol. The balanced equation is:  $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$ . After performing the reaction, the chemist isolates 30.09 grams of ammonia (NH<sub>3</sub>). Calculate the percentage yield of ammonia. 6 points total.



Tutor



Chat with the tutor to understand and answer the question.

Show My Answer

## Chat



TutorBot

Ah, dear student, it is wonderful to see your curiosity and dedication to learning chemistry. 🌸 Let's embark on this journey together, step by step, with patience and mindfulness.